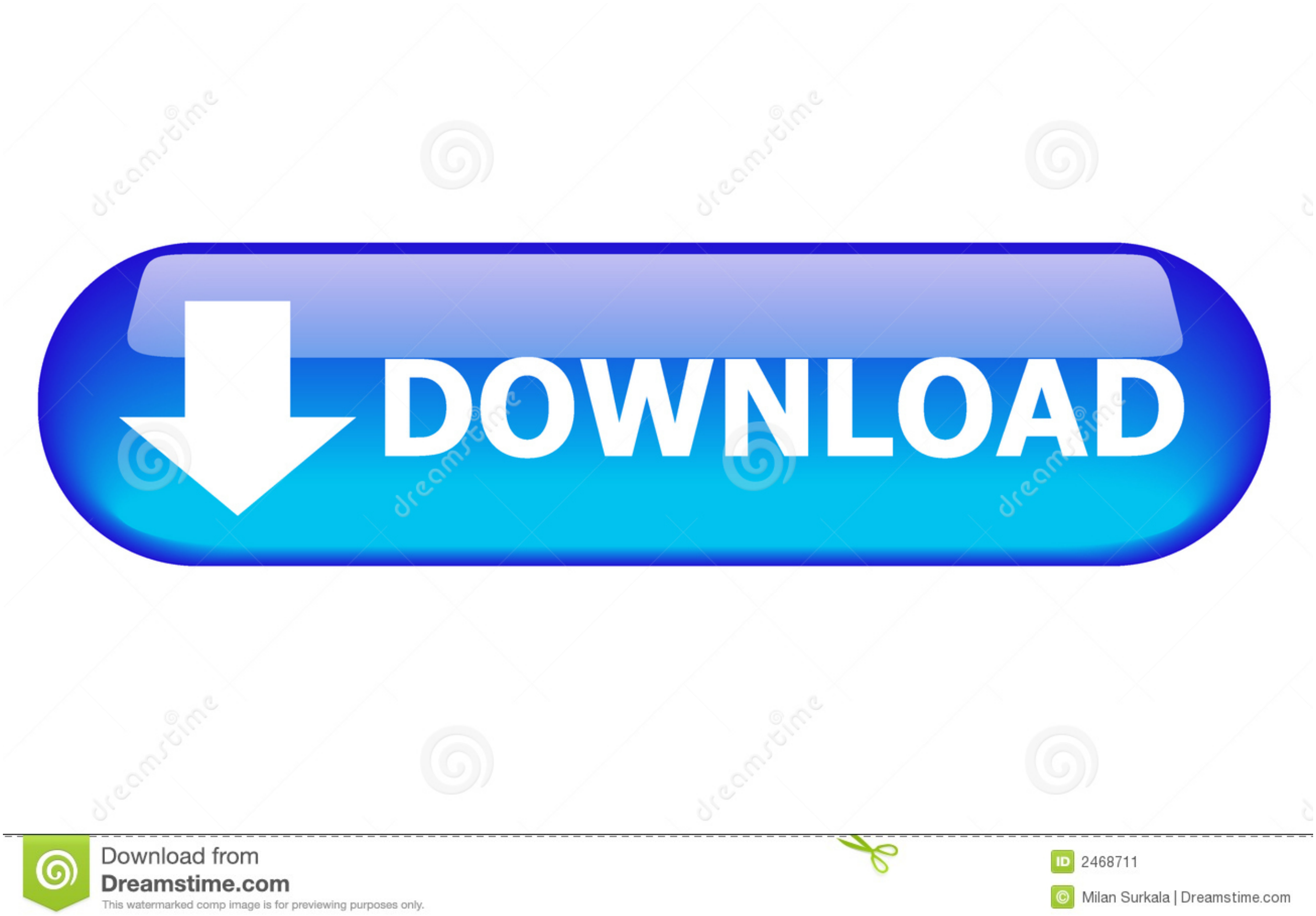


[Descargar Solucionario De Wangsness]

P19.18 (a) $n = \frac{PV}{RT} = \frac{(9.00 \text{ atm})(1.013\times10^5 \text{ Pa/atm})(8.00\times10^{-3} \text{ m}^3)}{(8.314 \text{ N-mol K})(293 \text{ K})} = \boxed{2.99 \text{ mol}}$

(b) $N = nN_A = (2.99 \text{ mol})(6.02\times10^{23} \text{ molecules/mol}) = \boxed{1.80\times10^{24} \text{ molecules}}$

P19.19 The equation of state of an ideal gas is $PV = nRT$ so we need to solve for the number of moles to find N.

$$n = \frac{PV}{RT} = \frac{(1.01\times10^5 \text{ N/m}^2)[(10.0 \text{ m})(20.0 \text{ m})(30.0 \text{ m})]}{(8.314 \text{ J/mol}\cdot\text{K})(293 \text{ K})} = 2.49\times10^5 \text{ mol}$$

$$N = nN_A = 2.49\times10^5 \text{ mol}(6.022\times10^{23} \text{ molecules/mol}) = \boxed{1.50\times10^{29} \text{ molecules}}$$

P19.20 $P = \frac{nRT}{V} = \left(\frac{9.00 \text{ g}}{18.0 \text{ g/mol}}\right)\left(\frac{8.314 \text{ J}}{\text{mol K}}\right)\left(\frac{773 \text{ K}}{2.00\times10^{-3} \text{ m}^3}\right) = \boxed{1.61 \text{ MPa}} = 15.9 \text{ atm}$

P19.21

$$\sum F_y = 0: \quad \rho_{out} gV - \rho_{in} gV - (200 \text{ kg})g = 0$$

$$(\rho_{out} - \rho_{in})(400 \text{ m}^3) = 200 \text{ kg}$$

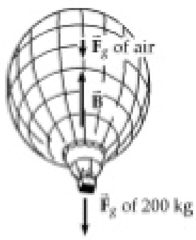


FIG. P19.21

The density of the air outside is 1.25 kg/m^3.

From $PV = nRT$, $\frac{n}{V} = \frac{P}{RT}$. This equation means that at constant pressure the density is inversely proportional to the temperature. Then the density of the hot air is

$$\rho_{in} = (1.25 \text{ kg/m}^3)\left(\frac{283 \text{ K}}{T_{in}}\right)$$

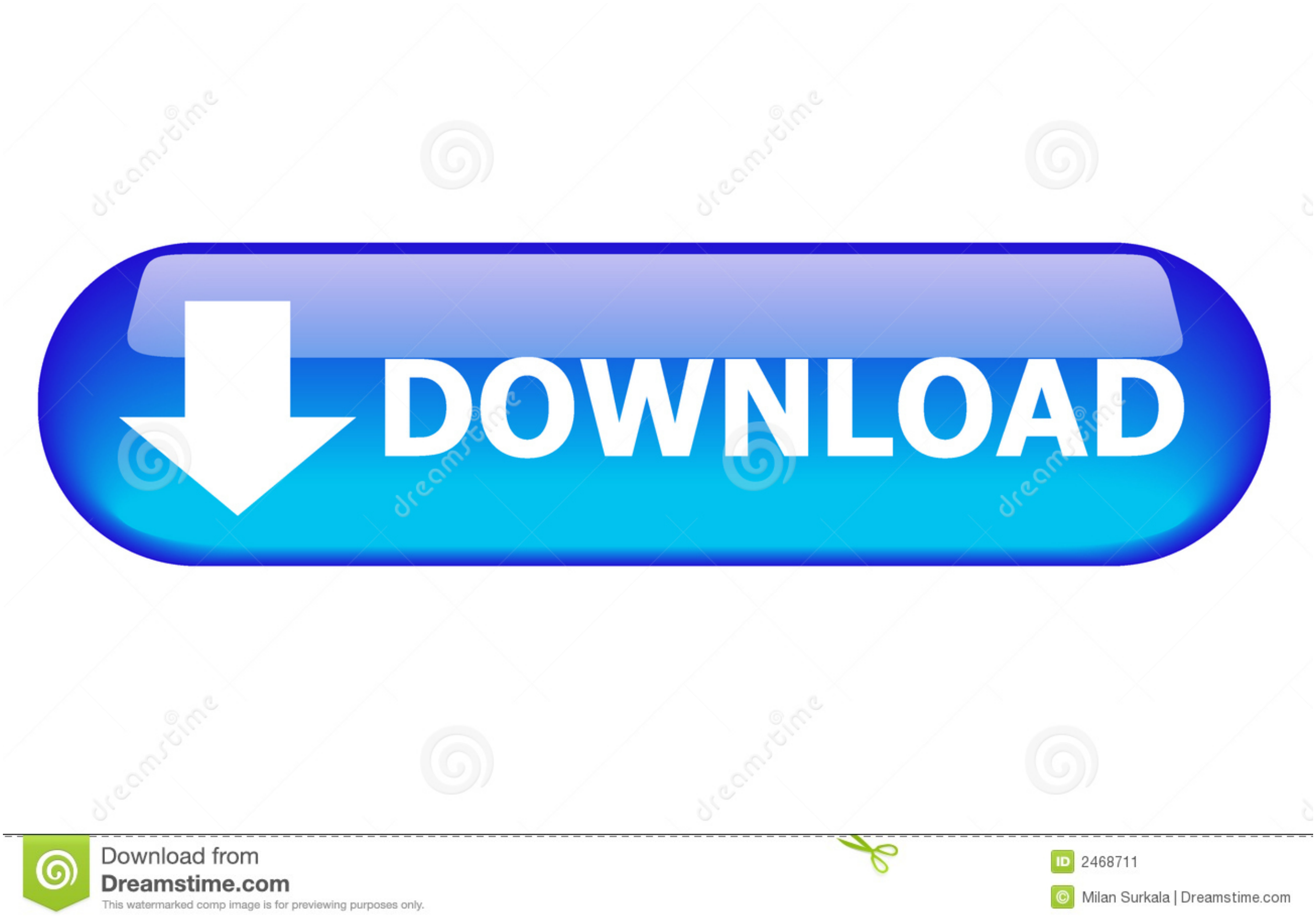
Then

$$(1.25 \text{ kg/m}^3)\left(1 - \frac{283 \text{ K}}{T_{in}}\right)(400 \text{ m}^3) = 200 \text{ kg}$$

$$1 - \frac{283 \text{ K}}{T_{in}} = 0.400$$

$$0.600 = \frac{283 \text{ K}}{T_{in}} \qquad T_{in} = \boxed{472 \text{ K}}$$

P19.22 Consider the air in the tank during one discharge process. We suppose that the process is slow enough that the temperature remains constant. Then as the pressure drops from 2.40 atm to 1.20 atm, the volume of the air doubles. During the first discharge, the air volume changes from 1 L to 2 L. Just 1 L of water is expelled and 3 L remains. In the second discharge, the air volume changes from 2 L to 4 L and 2 L of water is sprayed out. In the third discharge, only the last 1 L of water comes out. Each person could more efficiently use his device by starting with the tank half full of water.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